

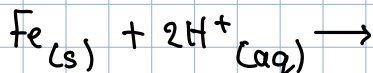
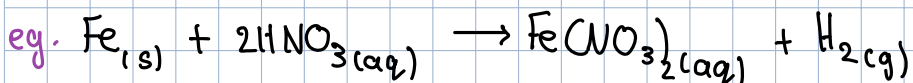
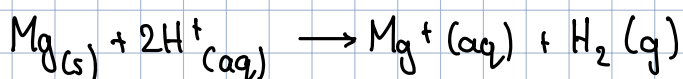
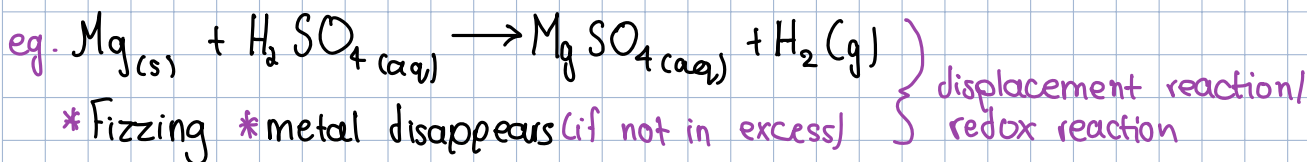
Typical reactions of acids

Reactions of acids

- Metals
 - Alkalis
 - Oxides
 - Carbohydrates
- } bases *alkalis are soluble in water

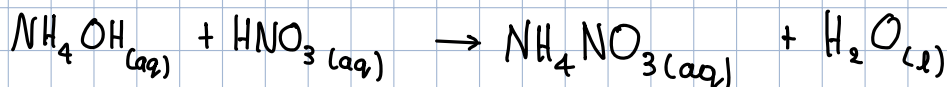
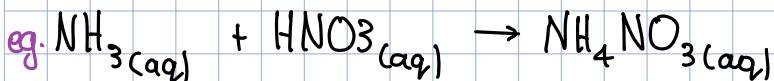
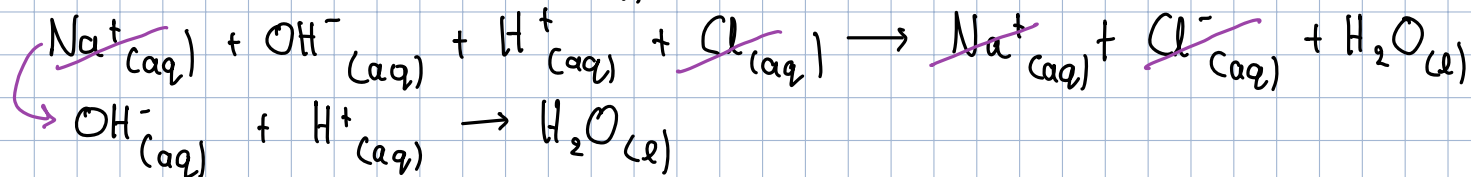
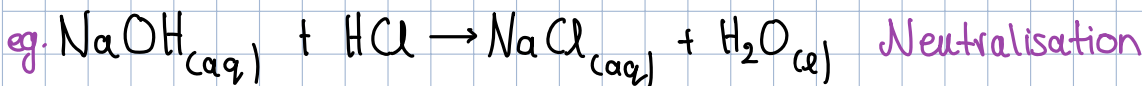
→ METALS

metal + acid → salt + hydrogen If the metal is more reactive than hydrogen



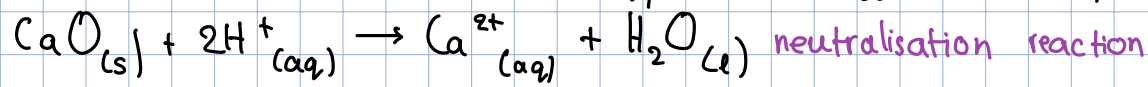
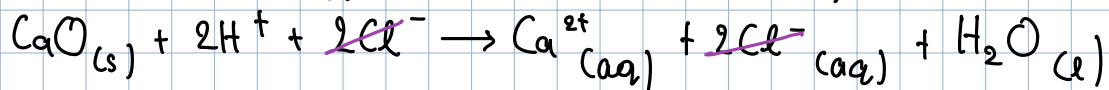
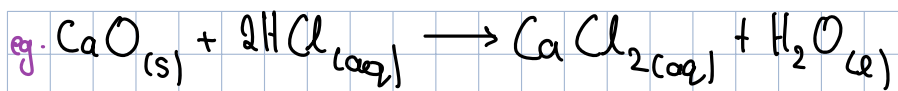
→ ALKALIS

alkali + acid → salt + water



→ METAL OXIDES

Metal oxides + Acid → Salt + Water * metal oxides are bases

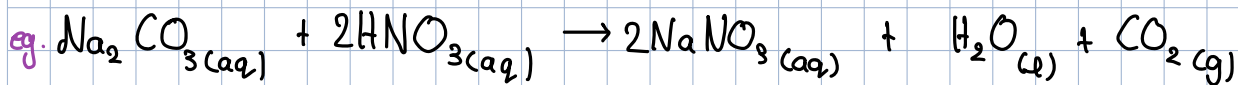


Important:

None of the neutralisation reactions are redox reactions since nothing gains or loses electrons and/or oxygen, ions are just being moved around. This means that it is not required for the metal to be above hydrogen in the reactivity series

→ CARBONATES

Carbonate + acid → Salt + water + carbon dioxide



* bubbles (main observation because of CO_2)

